

SOLUTIONS TO PROBLEMS

P1.2 Modeling the Earth as a sphere, we find its volume as

$$\frac{4}{3}\pi r^3 = \frac{4}{3}\pi(6.37 \times 10^6 \text{ m})^3 = 1.08 \times 10^{21} \text{ m}^3. \text{ Its density is then}$$

$$\rho = \frac{m}{V} = \frac{5.98 \times 10^{24} \text{ kg}}{1.08 \times 10^{21} \text{ m}^3} = \boxed{5.52 \times 10^3 \text{ kg/m}^3}. \text{ This value is intermediate between the tabulated densities of aluminum and iron. Typical rocks have densities around } 2\,000 \text{ to } 3\,000 \text{ kg/m}^3. \text{ The average density of the Earth is significantly higher, so higher-density material must be down below the surface.}$$

P1.17 Inserting the proper units for everything except G ,

$$\left[\frac{\text{kg m}}{\text{s}^2} \right] = \frac{G[\text{kg}]^2}{[\text{m}]^2}.$$

Multiply both sides by $[\text{m}]^2$ and divide by $[\text{kg}]^2$; the units of G are $\boxed{\frac{\text{m}^3}{\text{kg} \cdot \text{s}^2}}$.

P1.30
$$N_{\text{atoms}} = \frac{m_{\text{Sun}}}{m_{\text{atom}}} = \frac{1.99 \times 10^{30} \text{ kg}}{1.67 \times 10^{-27} \text{ kg}} = \boxed{1.19 \times 10^{57} \text{ atoms}}$$

P1.32

$$V = \frac{1}{3}Bh = \frac{[(13.0 \text{ acres})(43\,560 \text{ ft}^2/\text{acre})]}{3}(481 \text{ ft})$$

$$= 9.08 \times 10^7 \text{ ft}^3,$$

or

$$V = (9.08 \times 10^7 \text{ ft}^3) \left(\frac{2.83 \times 10^{-2} \text{ m}^3}{1 \text{ ft}^3} \right)$$

$$= \boxed{2.57 \times 10^6 \text{ m}^3}$$

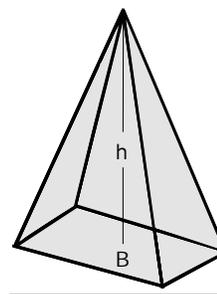


FIG. P1.32

P1.44 A typical raindrop is spherical and might have a radius of about 0.1 inch. Its volume is then approximately $4 \times 10^{-3} \text{ in}^3$. Since $1 \text{ acre} = 43\,560 \text{ ft}^2$, the volume of water required to cover it to a depth of 1 inch is

$$(1 \text{ acre})(1 \text{ inch}) = (1 \text{ acre} \cdot \text{in}) \left(\frac{43\,560 \text{ ft}^2}{1 \text{ acre}} \right) \left(\frac{144 \text{ in}^2}{1 \text{ ft}^2} \right) \approx 6.3 \times 10^6 \text{ in}^3.$$

The number of raindrops required is

$$n = \frac{\text{volume of water required}}{\text{volume of a single drop}} = \frac{6.3 \times 10^6 \text{ in}^3}{4 \times 10^{-3} \text{ in}^3} = 1.6 \times 10^9 \sim \boxed{10^9}.$$